Assignments must be submitted through email only to murgi@smail.iitm.ac.in. In parts which require hand writing, you can write neatly in a paper and take a photo/scan, use that in a document which is typed out for remaining parts. If two manuscripts are
 submitting 5 hrs late or more.

1. Consider air at $300 \mathrm{~K}, 1$ bar to be consisting of $\mathrm{N}_{2} \& \mathrm{O}_{2}$ alone, in ratio $79: 21$ by volume. If we take air at $300 \mathrm{~K}, 1 \mathrm{bar}$, and take it to various states given below, give the new equilibrium composition of the mixture. Consider only species $\mathrm{N}_{2}, \mathrm{O}_{2}, \mathrm{NO}, \mathrm{N}, \mathrm{O}$.
(a) $\mathrm{P}=1 \mathrm{bar}, \mathrm{T}=2500 \mathrm{~K}$;
(b) $\mathrm{P}=1 \mathrm{bar}, \mathrm{T}=6500 \mathrm{~K}$;
(c) $\mathrm{P}=0.25$ bar, $\mathrm{T}=2500 \mathrm{~K}$;
K ; (d) $\mathrm{P}=0.25$ bar, $\mathrm{T}=6500 \mathrm{~K}$;
(e) $\mathrm{P}=4$ bar, $\mathrm{T}=2500 \mathrm{~K}$; (f) $\mathrm{P}=4$ bar, $\mathrm{T}=6500 \mathrm{~K}$
(g) Check if your results satisfy the Le'Chatelier principle
2. Starting from Gibbs relation, derive an expression for Temperature in terms of Entropy. Use the tables for Ar at appropriate conditions, to test the values of entropy at two different temperatures.
3. Consider a gas whose molecules are flat discs sliding on a frictionless plane. (Eg. Carrom coins on a frictionless board). Assume that the molecules will collide similar to billiard balls, in this plane. Assume that the gas obeys $\mathrm{P}=\rho \mathrm{RT}$, where $\mathrm{P}=$ force per unit length of the edge of the board, $\rho=$ density per unit area.
(a) Find a relation between $\mathrm{P}, \rho$ and $\overline{C^{2}}$.
(b) Find by simple means, an expression for the mean free path.
(c) Using similar arguments used by Boltzmann for deriving the velocity distribution function, find the $f\left(c_{i}\right)$ for these new molecules.
4. A small satellite in the shape of a thin disc $(\mathrm{D}=50 \mathrm{~cm})$ is travelling through atmosphere of a remote planet. The satellite's diameter $\mathrm{D} \ll \lambda$, and it travels in a direction perpendicular to the disc surface with velocity ' +v '. The satellite's front surface is designed to be a sensor which can measure the collision frequency $\theta$ with which the molecules strike the satellite.
(a) If $\mathrm{v}=10 \mathrm{~km} / \mathrm{s}$, which is assumed to be far greater than the mean molecular speed in that atmosphere, and the measured value of $\theta=2.5 \mathrm{e}+21 / \mathrm{sec}$, find the number density of the molecules. Assume that the molecules reflect specularly from the surface of the disc. Check our $\mathrm{D} \ll \lambda$ assumption, using $\mathrm{d}=3.7 \mathrm{e}-10$ m.
(b) Find the drag on the satellite for the above case, if the mass of molecules is $3 \mathrm{e}-26 \mathrm{~kg}$.
(c) If $\mathrm{v}=1 \mathrm{~km} / \mathrm{s}$, which is comparable to mean speed of molecules, estimate the number density of the molecules using more detailed analysis from advanced kinetic theory. Use $\mathrm{T}=1000 \mathrm{~K}, \theta=2.5 \mathrm{e}+21 / \mathrm{sec}$, $\mathrm{m}=3 \mathrm{e}-26 \mathrm{~kg}$.
Note: the flux of molecules can be obtained from $\Gamma=\left(\pi \mathrm{D}^{2} / 4\right) \theta$, and molecules have comparable movement w.r.t. disc. Error function values can be evaluated using a commercial math software like Matlab. (d) Show that the drag coefficient on the satellite is given by:
$C_{D}=\frac{2}{s^{2}}\left\{\left(2 s^{2}+1\right) \operatorname{erf}(s)+\frac{2 s}{\sqrt{\pi}} e^{-s^{2}}\right\}$ where $s=\frac{v}{\sqrt{2 k T / m}}$ and $\operatorname{erf}(x)=\frac{2}{\sqrt{\pi}} \int_{0}^{x} e^{-z^{2}} d z$ Take into account the interaction of the molecules on both the surfaces of the disc.
(e) Find the net drag in the above case.
5. Consider the derivation in class for bimolecular collision rate, $\mathrm{Z}_{\mathrm{AB}}$. We used the limits $0-\infty$ for the dg integral. However, for reactions to occur, the relative velocity needs to have a value higher than a critical value such that $\varepsilon=1 / 2 \mathrm{mg}^{2}>\varepsilon_{a}$. So derive the expression for the rate of effective collisions between molecules A\&B. This will be the reaction rate expression used in chemical kinetics.
6. [BONUS] Consider the same plate as above, but this time with an angle of attack. Find the lift and drag coefficients in a similar manner as 4(d).
